

## Binary Ionic Compounds (Completion)

Name: \_\_\_\_\_ Per: \_\_\_\_\_

**Directions:** In the blank spaces, write the correct formula for the ionic compound made from the indicated cations and anions.

	Br <sup>-</sup>	S <sup>2-</sup>	N <sup>3-</sup>
Li <sup>+</sup>			
Sr <sup>2+</sup>			
Cu <sup>2+</sup>			
Al <sup>3+</sup>			
Pb <sup>4+</sup>			

**Directions:** First quickly scan the worksheet and circle any metals (as a symbol or as a name) that can have more than one charge. Then either give the name or chemical formula as necessary for each problem below.

1. LiF \_\_\_\_\_

11. lithium chloride \_\_\_\_\_

2. Li<sub>2</sub>O \_\_\_\_\_

12. lithium nitride \_\_\_\_\_

3. AlCl<sub>3</sub> \_\_\_\_\_

13. aluminum fluoride \_\_\_\_\_

4. BeO \_\_\_\_\_

14. magnesium oxide \_\_\_\_\_

5. CuF \_\_\_\_\_

15. calcium chloride \_\_\_\_\_

6. Cu<sub>2</sub>O \_\_\_\_\_

16. copper (II) oxide \_\_\_\_\_

7. Cu<sub>3</sub>N \_\_\_\_\_

17. copper (II) nitride \_\_\_\_\_

8. Na<sub>2</sub>O \_\_\_\_\_

18. sodium nitride \_\_\_\_\_

9. FeP \_\_\_\_\_

19. lead (IV) fluoride \_\_\_\_\_

10. PbS \_\_\_\_\_

20. lead (IV) sulfide \_\_\_\_\_