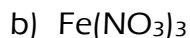
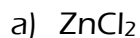


Chapter 10 Mock Test

Part 1: Molar Mass – Determine the molar mass for the following compounds. Round to the nearest hundredths and remember your units!



Part 2: Conversions between Mass, Volume, & Representative Particles – Complete the following questions/conversions. SHOW ALL WORK. Do you forget correct units. Round to the nearest hundredths.

1. One mole of any element/compound/ion always equals _____ representative units.
2. One mole of any gas at STP is equal to _____ liters.
3. How many molecules are in 8.17 L of HCl at STP?

4. How many formula units are in 94 g of NaCl?

5. What is the volume of 325 grams of Krypton at STP?

6. How many atoms are in 5.48 L of CO_2 at STP?

2. A compound contains 75.46% carbon, 4.43% hydrogen, and 20.11% oxygen by mass. It has a molar weight of 318.31 g/mol.

○ The empirical formula of the compound is _____.

○ The molecular formula of this compound is _____.