Monoatomic & Polyatomic Ionic Compounds (Points) Name:					
Directions: In the blank spaces, write the correct formula for the ionic compound made from the indicated cations and anions.					
	S ²⁻	Br-1	SO ₄ ² -	PO ₄ ³⁻	
Rb+					
NH ₄ ⁺					
Ca ²⁺					
Fe ³⁺					
Pb ⁴⁺					
Directions: First quickly scan the worksheet and circle any metals (as a symbol or as a name) that are transition metals or can form more than one ion. Then either give the name or chemical formula as necessary for each problem below.					
1. FeCl ₃		11. C	opper (II) chloride _		
2. K ₂ SO ₄		12. C	alcium phosphate _		
3. Pb(OH) ₂		13. Lit	13. Lithium oxide		
4. Al ₂ O ₃	4. Al ₂ O ₃		14. Copper (II) nitrite		
5. AIPO ₃	AIPO ₃		15. Cadmium nitrate		
6. NH ₄ Br		16. Tir	n (IV) phosphite		
7. Sn(CN) ₄		17. Zir	nc cyanide		
8. Na ₂ CO ₃		18. lrc	on (II) sulfide		
9. MnO ₂	9. MnO ₂		19. Barium nitride		
10. CoF ₂		20. Pc	20. Potassium iodide		