

More Practice Percent Composition (Points)

Instructions: Calculate the percent composition of each element in the following compounds. Round your answer to the nearest tenth of a decimal place. Circle your final answer. SHOW ALL WORK.

Percent Composition from Mass

1) A compound was decomposed into its elements and gave 5.217 grams of carbon, 0.9620 grams of hydrogen, and 2.478 grams of oxygen. What is the percentage composition of each element in the compound

2) A compound contains 7.20 g of C, 2.20 g of hydrogen and 17.6 g of oxygen. Calculate the percentage composition for the elements in the compound.

3) Find the percent composition of each element for a compound containing carbon and oxygen if 30.80 g of the compound contains 8.40 grams of carbon.

Percent Composition from Chemical Formula

4) What is the percent composition for each element in sodium chloride, NaCl?

5) What is the percent composition for each element in Phosphoric acid, H₃PO₄?

Percent Composition as a Conversion Factor

Instructions: Calculate the mass of the element in the given mass of compound using percent composition. Round your answer to the nearest tenth of a decimal place. Circle your final answer. SHOW ALL WORK.

6) Mass of chlorine in 350g of Cl₂O₇

7) Mass of hydrogen in 120g of H₂O₂.